Province of the
EASTERN CAPE
EDUCATION

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ERRATUM

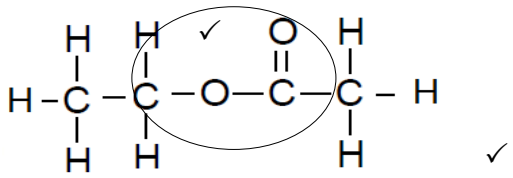
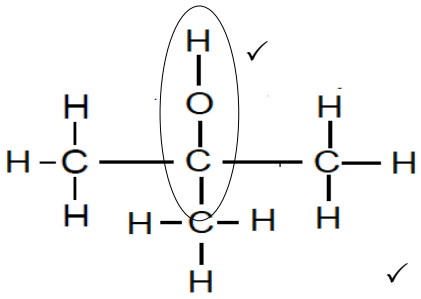
TO: **CHIEF EDUCATION SPECIALISTS**
DISTRICT CURRICULUM COORDINATORS
DISTRICT ASSESSMENT OFFICIALS (DAOs
DISTRICT SUBJECT ADVISORS (DSAs)
PROVINCIAL SUBJECT COORDINATORS
CIRCUIT MANAGERS
DEPUTY CHIEF EDUCATION SPECIALISTS
SENIOR EDUCATION SPECIALISTS
PRINCIPALS OF SCHOOLS IN THE FET BAND

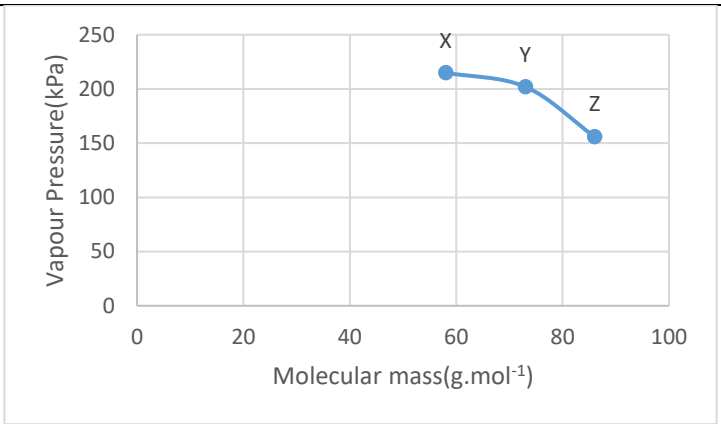
SUBJECT: **ERRATUM – TECHNICAL SCIENCE PAPER 2**

DATE: **11 JUNE 2024**

CORRECTED RESPONSE FOR TECHNICAL SCIENCES PAPER 2

QUESTION 2		
2.1	An atom or a group of atom that determines the chemistry of a molecule.	(2)
2.2.2	A or D	
2.2.4	A /D/H	
2.2.5	**Question cancelled as there are no isomers in the given options.**	
2.3.1	Pent-1-ene or 1-pentene	

2.4.1	 ✓ ✓ **MARKING CRITERIA :** • Correct functional group ✓ • Whole structure correct ✓ • **NOTE: If a bond or hydrogen is missing 1/2**	
2.4.3	C_6H_{12}	
2.4.4	<u>Carbonyl</u> group	
2.4.5	 ✓ ✓ **MARKING CRITERIA :** • Correct functional group ✓ • *Whole* structure correct ✓ • **NOTE: If a bond or hydrogen is missing 1/2**	

QUESTION 3		
3.1	The pressure exerted by a vapour at equilibrium with its liquid in a closed system.	
3.2		
3.3	The greater the molecular mass of the organic compound/alkanes, the lower the vapour pressure.	
	OR	
	As the molecular mass increases ,vapour pressure decreases.	
3.5	• Referring to chain/molecular mass. • Comparison of strength of intermolecular force. • Intermolecular forces in compound Z are stronger than that of Compound Y. • More energy is needed to overcome the intermolecular forces in D than in Y.	
3.6	Z	
QUESTION 4		
4.1	Mark allocation 2 or 0	

4.3	If a learner writes London forces ONLY, no mark. Dipole-dipole (intermolecular forces)	
4.4	Boiling point increases with an increase in number of chlorine atoms/Boiling point increases with an increase in molecular mass.	
4.5.2	One independent variable	
4.5.3	Vapour pressure of compound I will be lower than the vapour pressure of Compound L.	
QUESTION 6		
6.3	Positive holes/Holes that are positively charged.	
6.6	Improves conductivity (of the semiconductor.	

We sincerely apologise for any inconvenience we might have caused.

Yours in education.

MRS P.E. JAPHTA 11 June 2024
(A) CES: AIDIBM SUBDIRECTORATE **DATE**