



Province of the  
**EASTERN CAPE**  
EDUCATION

Iphondo leMpuma Kapa: Isebe leMfundo  
Provinsie van die Oos Kaap: Departement van Onderwys  
Porafensie Ya Kapa Botjhabela: Lefapha la Thuto

# **NATIONAL SENIOR CERTIFICATE**

## **GRADE 12**

### **JUNE 2025**

## **PHYSICAL SCIENCES: (CHEMISTRY) P2**

**MARKS: 150**

**TIME: 3 hours**



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This question paper consists of 18 pages, including 2 data sheets.

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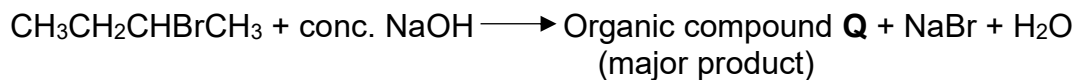
**INSTRUCTIONS AND INFORMATION**

1. Write your name and surname in the appropriate space on the ANSWER BOOK.
2. This question paper consists of SEVEN questions. Answer ALL the questions in the ANSWER BOOK.
3. Start EACH question on a NEW page in the ANSWER BOOK.
4. Number the answers correctly according to the numbering system used in this question paper.
5. Leave ONE line between two sub-questions, for example between QUESTION 2.1 and QUESTION 2.2.
6. You may use a non-programmable calculator.
7. You may use appropriate mathematical instruments.
8. Show ALL formulae and substitutions in ALL calculations.
9. Round off your FINAL numerical answers to a minimum of TWO decimal places.
10. Give brief motivations, discussions, et cetera where required.
11. You are advised to use the attached DATA SHEETS.
12. Write neatly and legibly.

**QUESTION 1: MULTIPLE-CHOICE QUESTIONS**

Various options are provided as possible answers to the following questions. Choose the answer and write only the letter (A–D) next to the question numbers (1.1 to 1.10) in the ANSWER BOOK, for example 1.11 E.

- 1.1 Which ONE of the following compounds is an alkene?
- A  $\text{C}_3\text{H}_8$
- B  $\text{C}_3\text{H}_6$
- C  $\text{C}_3\text{H}_4$
- D  $\text{C}_3\text{H}_6\text{O}$  (2)
- 1.2 Which ONE of the following is NOT correct about compounds that belong to the same homologous series?
- A Similar chemical properties
- B They have the same general formula
- C They have the same functional group
- D Similar physical properties (2)
- 1.3 Which ONE of the following can form a TERTIARY ALCOHOL?
- A  $\text{CH}_3\text{OH}$
- B  $\text{C}_2\text{H}_5\text{OH}$
- C  $\text{C}_3\text{H}_7\text{OH}$
- D  $\text{C}_4\text{H}_9\text{OH}$  (2)
- 1.4 Consider the reaction below:



Which ONE of the following CORRECTLY gives the type of reaction and the name of organic compound Q?

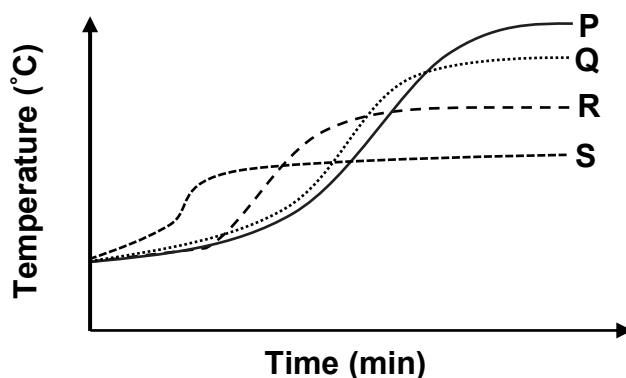
	TYPE OF REACTION	ORGANIC COMPOUND Q
A	Elimination	But-1-ene
B	Elimination	But-2-ene
C	Addition	But-1-ene
D	Addition	But-2-ene

(2)

1.5 Consider the four organic compounds below:

butan-1-ol, propanoic acid, 2-methylpropanal and butanal

The heating curves for the four organic compounds were obtained.



Which curve represents the heating curve of butanal?

A Curve **P**

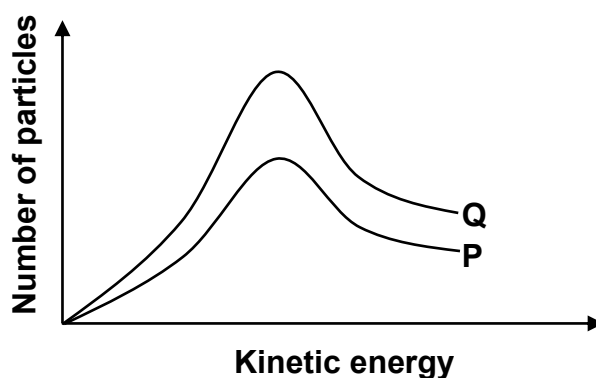
B Curve **Q**

C Curve **R**

D Curve **D**

(2)

1.6 The Maxwell-Boltzmann energy distribution curve **P** for  $\text{CO}_2$  gas under certain conditions. Curve **Q** was obtained after a change was made.



Which ONE of the following represents the change made to obtain curve **Q**?

A Increase in temperature

B Addition of a catalyst

C Increase in concentration

D Increase in surface area of  $\text{CO}_2$

(2)

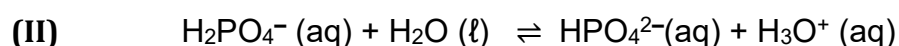
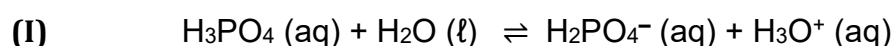
- 1.7 Consider the following reversible hypothetical reaction:



The heat of reactants ( $H_r$ ) for the forward reaction is  $25 \text{ kJ}\cdot\text{mol}^{-1}$  and activation energies ( $E_a$ ) for the forward reaction and reverse reaction is  $35 \text{ kJ}\cdot\text{mol}^{-1}$  and  $45 \text{ kJ}\cdot\text{mol}^{-1}$  respectively.

The heat of the products ( $H_p$ ) for the forward reaction is ...

- A  $10 \text{ kJ}\cdot\text{mol}^{-1}$ .
- B  $15 \text{ kJ}\cdot\text{mol}^{-1}$ .
- C  $20 \text{ kJ}\cdot\text{mol}^{-1}$ .
- D  $35 \text{ kJ}\cdot\text{mol}^{-1}$ . (2)
- 1.8 Which ONE of the following acids, with the same concentration, will have the highest conductivity at a given temperature?
- A  $\text{H}_2\text{CO}_3$
- B  $\text{CH}_3\text{COOH}$
- C  $\text{HCl}$
- D  $\text{H}_2\text{SO}_4$  (2)
- 1.9 Consider the two-step ionisation of phosphoric acid:



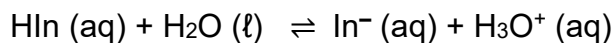
Consider the following statements regarding the two-step ionisation of  $\text{H}_3\text{PO}_4$ .

- I  $\text{H}_2\text{PO}_4^-$  act as an ampholyte
- II  $\text{HPO}_4^{2-}$  is the conjugate acid of  $\text{H}_2\text{PO}_4^-$
- III  $\text{H}_2\text{PO}_4^-$  can be considered as a diprotic acid

Which ONE of the above statement(s) is/are TRUE?

- A I only
- B I and II only
- C I and III only
- D II and III only (2)

- 1.10 A specific indicator is colourless in an acidic solution and pink in an alkaline solution. The general equation for the indicator is:



The indicator is added to a sodium hydroxide (NaOH) solution.

Which ONE of the following combinations are CORRECT regarding the colour of HIn and In<sup>-</sup> and the shift in the equilibrium position?

	HIn	In <sup>-</sup>	SHIFT IN EQUILIBRIUM POSITION
A	Pink	Colourless	Right
B	Colourless	Pink	Right
C	Pink	Colourless	Left
D	Colourless	Pink	Left

(2)  
[20]

**QUESTION 2 (Start on a new page.)**

The table below shows organic molecules (**A** to **E**) from different homologous series.

<b>A</b>	2-methylpentan-3-one	<b>B</b>	$  \begin{array}{c}  \text{O} \\  \parallel \\  \text{CH}_3\text{CH} - \text{CH} - \text{C} - \text{H} \\    \\  \text{CH}_3  \end{array}  $
<b>C</b>	$\text{CH}_3\text{C}\equiv\text{CCH}(\text{CH}_3)\text{CH}_3$	<b>D</b>	$  \begin{array}{c}  \text{Br} \quad \text{CH}_2\text{CH}_3 \\    \quad   \\  \text{CH}_3\text{CH} - \text{C} - \text{C} - \text{H} \\    \quad   \quad   \\  \text{Br} \quad \text{H} \quad \text{CH}_2\text{CH}_3  \end{array}  $
<b>E</b>	Ethyl butanoate		

- 2.1 Define *functional group*. (2)
- 2.2 Write down the LETTER of the organic compound that represents the following:
- 2.2.1 Carbonyl group bonded to two carbon atoms (1)
- 2.2.2 Has the general formula  $\text{C}_n\text{H}_{2n-2}$  (1)
- 2.2.3 Is an aldehyde (1)
- 2.3 Write down the IUPAC name of:
- 2.3.1 Compound **B** (2)
- 2.3.2 Compound **C** (2)
- 2.3.3 Compound **D** (3)
- 2.4 Write down the:
- 2.4.1 STRUCTURAL FORMULA of compound **A** (2)
- 2.4.2 Name of the reaction that occurred to produce compound **E** (1)
- 2.4.3 STRUCTURAL FORMULA of the carboxylic acid needed to produce compound **E** (2)

2.5 An unknown organic compound ( $C_xH_yO_z$ ) with a molar mass of  $74 \text{ g}\cdot\text{mol}^{-1}$  consists of 43,24% oxygen by mass.

2.5.1 Determine the MOLECULAR FORMULA of the organic compound. (3)

2.5.2 Draw TWO STRUCTURAL FORMULAE for the functional isomers that are represented by the molecular formula in QUESTION 2.5.1. (4)

**[24]**



**QUESTION 3 (Start on a new page.)**

Learners investigate the effect of structural differences on the boiling points of straight-chain PRIMARY ALCOHOLS. The data from the investigation are shown in the table below.

NUMBER OF CARBON ATOMS	BOILING POINTS OF THE ALCOHOLS (20 °C)
1	64
2	78
3	98
4	118

- 3.1 Define *boiling point*. (2)
- 3.2 Write down the controlled variable for this investigation. (1)
- 3.3 Explain the trend observed in the boiling points of the alcohols. (4)
- 3.4 Will the vapour pressure of the alcohols INCREASE, DECREASE or REMAIN THE SAME with an increase in the number of carbon atoms?  
Give a reason for the answer. (2)

A PRIMARY ALCOHOL has a boiling point of 108 °C.

- 3.5 Give the IUPAC name of the primary alcohol with a boiling point of 108 °C. (2)
- 3.6 Fully explain the answer to QUESTION 3.5. (4)

**[15]**

**QUESTION 4 (Start on a new page.)**

- 4.1 Heptane undergoes cracking to produce a FIVE CARBON alkane **P** and organic compound **Q**, as shown below.



4.1.1 Define *cracking*. (2)

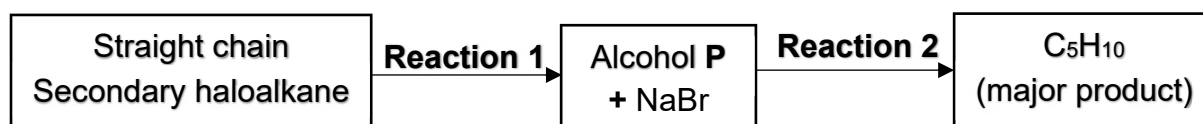
4.1.2 Write down the MOLECULAR FORMULA for compound **Q**. (2)

Compound **P** undergoes complete combustion.

4.1.3 Write down the most important use of alkanes. (1)

4.1.4 Using MOLECULAR FORMULAE, write down balanced equation for the complete combustion for COMPOUND **P**. (3)

- 4.2 Consider the flow diagram showing organic reactions given below.



Consider **REACTION 1**.

Write down:

4.2.1 The IUPAC name of the SECONDARY HALOALKANE (2)

4.2.2 The name of the reaction (1)

4.2.3 One reaction condition besides heat (1)

4.2.4 The STRUCTURAL FORMULA of alcohol **P** (2)

Consider **REACTION 2**.

Write down the:

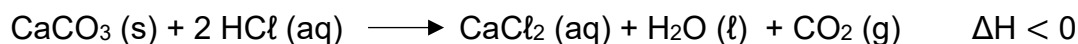
4.2.5 Name of the type of elimination reaction (1)

4.2.6 STRUCTURAL FORMULA for the product that was produced (2)

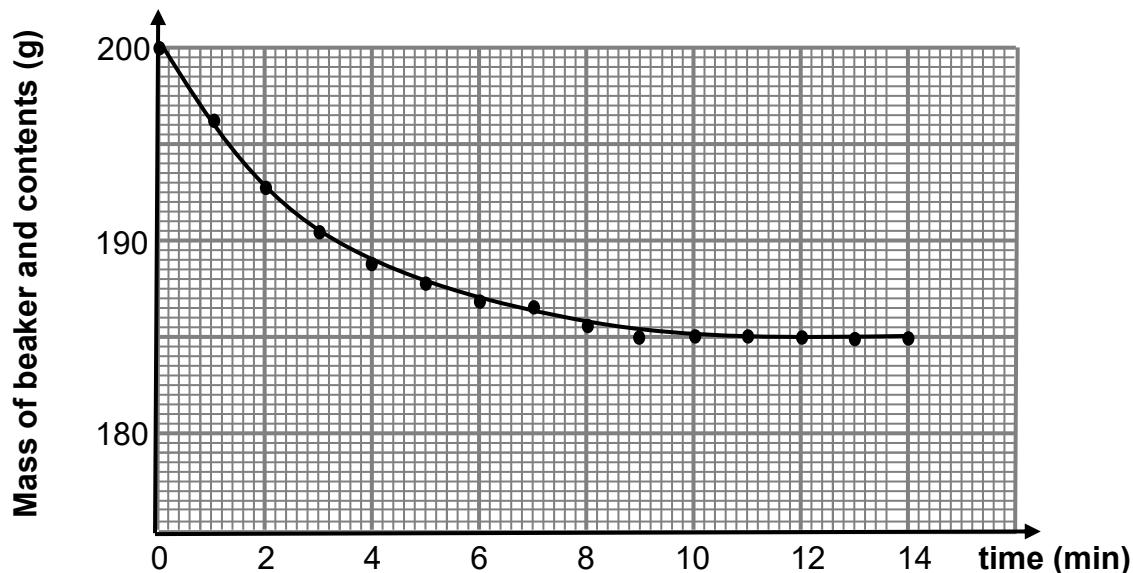
[17]

**QUESTION 5 (Start on a new page.)**

Calcium carbonate ( $\text{CaCO}_3$ ) chunks are added to EXCESS dilute hydrochloric acid ( $\text{HCl}$ ) solution in an Erlenmeyer flask that is placed on an electronic scale, as shown below. The balanced equation for the reaction that takes place is:

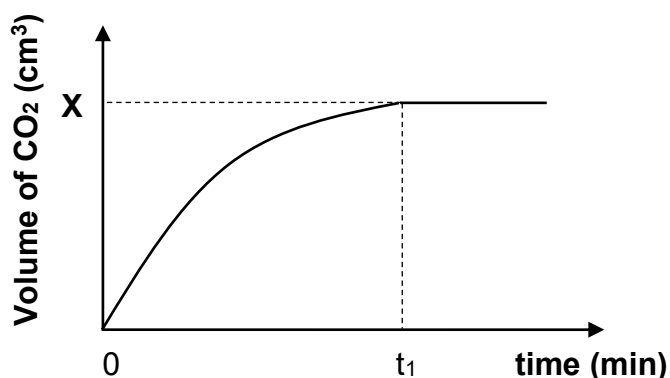


The change in mass of the flask and its contents are recorded in 1-minute intervals. The results obtained are shown in the graph below.



- 5.1 Is the reaction EXOTHERMIC or ENDOTHERMIC? Give a reason for the answer. (2)
- 5.2 Give a reason why the mass of the flask and its contents does not remain constant. (2)
- 5.3 Write down the reading on the scale balance on 4 minutes. (1)

- 5.4 The sketch graph below shows the change in volume of  $\text{CO}_2$  gas produced for this experiment.



- 5.4.1 Write down the value of  $t_1$ . (1)

- 5.4.2 How does the rate at which the amount  $\text{CO}_2$  is produced compared to the rate at which the amount  $\text{CaCO}_3$  is consumed?

Choose from HIGHER THAN, SMALLER THAN or EQUAL TO.

Give a reason for the answer. (3)

Calculate the:

- 5.4.3 Value of **X**  
Take the molar volume as  $24\,000\text{ cm}^3\cdot\text{mol}^{-1}$  (5)

- 5.4.4 Average rate in  $\text{g}\cdot\text{min}^{-1}$  at which calcium carbonate was consumed after 11 minutes (4)

- 5.5 The experiment is repeated by increasing the temperature of the reaction mixture. The results of the two experiments are compared.

- 5.5.1 Write down a hypothesis for this comparison. (2)

- 5.5.2 How would reaction rate be affected by this change?

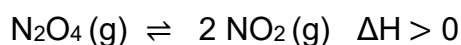
Choose from INCREASES, DECREASES or REMAINS THE SAME. (1)

- 5.5.3 Explain the answer to QUESTION 5.5.2 by referring to the collision theory. (3)

**[24]**

**QUESTION 6 (Start on a new page.)**

- 6.1 Dinitrogen tetroxide ( $\text{N}_2\text{O}_4$ ) decomposes to nitrogen dioxide ( $\text{NO}_2$ ) according to the balanced equation:



The reaction is allowed to reach chemical equilibrium.

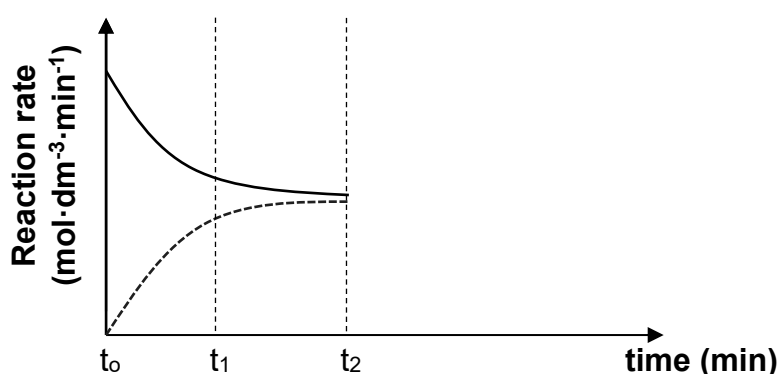
- 6.1.1 State Le Chatelier's principle in words. (2)

- 6.1.2 How would EACH of the following changes effect the concentration of  $\text{NO}_2$  at equilibrium.

Choose from INCREASES, DECREASES or NO EFFECT.

- (a) Addition of a suitable indicator (1)
- (b) Increase in temperature (1)
- (c) Increase in pressure by decreasing the volume. (4)  
Explain the answer by referring to Le Chatelier's principle.

- 6.2 The reaction rate-time graph below shows the reaction until equilibrium.



- 6.2.1 Write down the reaction represented by the dashed line. (2)

- 6.2.2 The concentration of  $\text{NO}_2$  was increased at  $t_2$ .

Redraw the graph above in the ANSWER BOOK on the same set of axes, sketch the complete graph showing the effect of the increase in the concentration of  $\text{NO}_2$  after  $t_2$ . (2)

- 6.3 The table below shows the experimental data for the  $\text{NO}_2\text{--N}_2\text{O}_4$  system at  $25^\circ\text{C}$ .

INITIAL CONCENTRATIONS ( $\text{mol}\cdot\text{dm}^{-3}$ )		EQUILIBRIUM CONCENTRATIONS ( $\text{mol}\cdot\text{dm}^{-3}$ )	
$[\text{NO}_2]$	$[\text{N}_2\text{O}_4]$	$[\text{NO}_2]$	$[\text{N}_2\text{O}_4]$
0,05	0,446	0,0457	0,448

- 6.3.1 Calculate the equilibrium constant at  $25^\circ\text{C}$ . (3)

- 6.3.2 Is there a HIGH YIELD or LOW YIELD of  $\text{NO}_2$  at  $25^\circ\text{C}$ .

Give a reason for the answer. (2)

When the initial concentration of  $\text{NO}_2$  was changed to  $0,03\text{ mol}\cdot\text{dm}^{-3}$ , it is found that the equilibrium concentration of  $\text{N}_2\text{O}_4$  is now  $0,491\text{ mol}\cdot\text{dm}^{-3}$  at  $25^\circ\text{C}$ .

- 6.3.3 Calculate the percentage decomposition of  $\text{N}_2\text{O}_4$  when the concentration of  $\text{NO}_2$  was changed to  $0,03\text{ mol}\cdot\text{dm}^{-3}$  at  $25^\circ\text{C}$ . (6)
- [23]**

**QUESTION 7 (Start on a new page.)**

7.1 The table below shows the ionisation constants,  $K_a$  values, for three acids at 25 °C.

NAME	FORMULA	$K_a$ value
Methanoic acid	HCOOH	$1,8 \times 10^{-4}$
Ethanoic acid	CH <sub>3</sub> COOH	$1,8 \times 10^{-5}$
Propanoic acid	CH <sub>3</sub> CH <sub>2</sub> COOH	$1,3 \times 10^{-5}$

7.1.1 Define an *acid* according to the Lowry-Brønsted theory. (2)

7.1.2 Write down the ionisation reaction for CH<sub>3</sub>COOH. (2)

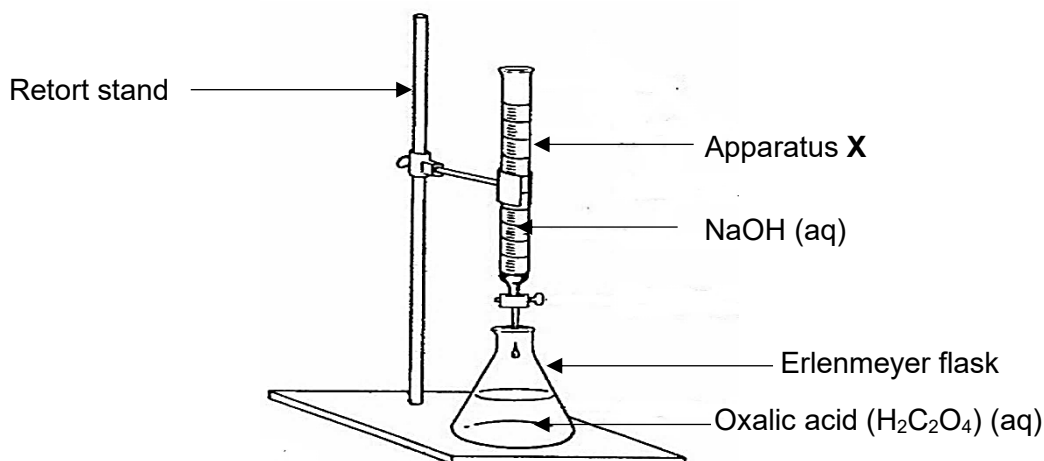
7.1.3 Are the above WEAK ACIDS or STRONG ACIDS?

Give a reason for the answer. (2)

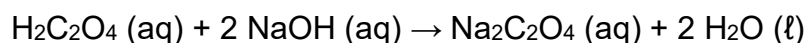
7.1.4 Which ONE of the three acids, with equal concentrations, will have the lowest pH value?

Explain the answer. (3)

7.2 A group of learners used the set-up below to titrate sodium hydroxide (NaOH) against oxalic acid (H<sub>2</sub>C<sub>2</sub>O<sub>4</sub>).



The balanced equation is:



7.2.1 Write down the name of apparatus X. (1)

7.2.2 Give a reason why the titration is carried out at least three times. (1)

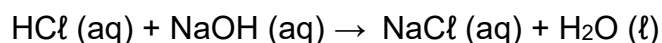
7.2.3 At what pH range will a suitable indicator changes colour for this titration?

Choose from:

3,1– 4,4	6,0–7,6	8,3–10	(1)
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7.2.4 Explain the answer to QUESTION 7.2.3 by referring to the relevant equation. (3)

7.3 The learners carried out another titration between hydrochloric acid (HCl) and sodium hydroxide (NaOH) with EQUAL concentration. They placed 25 cm<sup>3</sup> of hydrochloric acid with a concentration of 0,1 mol·dm<sup>-3</sup> in the Erlenmeyer flask. The balanced equation is:



They over titrated the sodium hydroxide solution. The pH value at the end is 12,52.

Calculate the:

7.3.1 Initial number of moles of hydrochloric acid (3)

7.3.2 Concentration of the excess hydroxide ions (4)

7.3.3 Volume of sodium hydroxide titrated in cm<sup>3</sup> (5)

[27]

**TOTAL: 150**



**NATIONAL SENIOR CERTIFICATE  
NASIONALE SENIOR SERTIFIKAAT**

**DATA FOR PHYSICAL SCIENCES GRADE 12  
PAPER 2 (CHEMISTRY)**

**GEGEWENS VIR FISIESTE WETENSKAPPE GRAAD 12  
VRAESTEL 2 (CHEMIE)**

**TABLE 1: PHYSICAL CONSTANTS/TABEL 1: FISIESTE KONSTANTES**

NAME/NAAM	SYMBOL/SIMBOOL	VALUE/WAARDE
Standard pressure <i>Standaarddruk</i>	$p^\theta$	$1,013 \times 10^5 \text{ Pa}$
Molar gas volume at STP <i>Molêre gasvolume teen STD</i>	$V_m$	$22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$
Standard temperature <i>Standaardtemperatuur</i>	$T^\theta$	$273 \text{ K}$
Charge on electron <i>Lading op elektron</i>	$e$	$-1,6 \times 10^{-19} \text{ C}$
Avogadro's constant <i>Avogadro se konstante</i>	$N_A$	$6,02 \times 10^{23} \text{ mol}^{-1}$

**TABLE 2: FORMULAE/TABEL 2: FORMULES**

$n = \frac{m}{M}$ or/of $n = \frac{N}{N_A}$ or/of $n = \frac{V}{V_m}$	$c = \frac{n}{V}$ or/of $c = \frac{m}{MV}$ $\frac{c_a V_a}{c_b V_b} = \frac{n_a}{n_b}$	$\text{pH} = -\log[\text{H}_3\text{O}^+]$ $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1 \times 10^{-14}$ at/by 298 K
-----------------------------------------------------------------------------	-------------------------------------------------------------------------------------------	-----------------------------------------------------------------------------------------------------------------------------

TABLE 3: THE PERIODIC TABLE OF ELEMENTS/TABEL 3: DIE PERIODIEKE TABEL VAN ELEMENTE

KEY/ SLEUTEL																		Atomgetal Atomic number		Simbools Symbol									
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18												
(I)		(II)																(III)	(IV)	(V)	(VI)	(VII)	(VIII)						
1 H 1																		5 B 11	6 C 12	7 N 14	8 O 16	9 F 19	10 Ne 20						
3 Li 7	4 Be 9																13 Al 27	14 Si 28	15 P 31	16 S 32	17 Cl 35,5	18 Ar 40							
11 Na 23	12 Mg 24																31 Ga 70	32 Ge 73	33 As 75	34 Se 79	35 Br 80	36 Kr 84							
19 K 39	20 Ca 40	21 Sc 45	22 Ti 48	23 V 51	24 Cr 52	25 Mn 55	26 Fe 56	27 Co 59	28 Ni 59	29 Cu 63,5	30 Zn 65							49 In 115	50 Sn 119	51 Sb 122	52 Te 128	53 I 127	54 Xe 131						
37 Rb 86	38 Sr 88	39 Y 89	40 Zr 91	41 Nb 92	42 Mo 96	43 Tc 98	44 Ru 101	45 Rh 103	46 Pd 106	47 Ag 108	48 Cd 112	81 Tl 204	82 Pb 207	83 Bi 209	84 Po 210	85 At 210	86 Rn 222												
55 Cs 133	56 Ba 137	57 La 139	72 Hf 179	73 Ta 181	74 W 184	75 Re 186	76 Os 190	77 Ir 192	78 Pt 195	79 Au 197	80 Hg 201							115 At 210	116 Po 210	117 Bi 209	118 Og 294	119 Ts 294	120 Og 294						
87 Fr 223	88 Ra 226	89 Ac																											

58 Ce 140	59 Pr 141	60 Nd 144	61 Pm	62 Sm 150	63 Eu 152	64 Gd 157	65 Tb 159	66 Dy 163	67 Ho 165	68 Er 167	69 Tm 169	70 Yb 173	71 Lu 175
90 Th 232	91 Pa 231	92 U 238	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr











